INDIAN SCHOOL MUSCAT
HALF YEARLY EXAMINATION
SEPTEMBER 2019

SET C

CLASS X

Marking Scheme – SUBJECT[CHEMISTRY][THEORY]

Q.NO.	Answers	Marks (with split up)
1.	PHY- MCQ	
2.	PHY- MCQ	
3.	PHY- MCQ	
4.	PHY- MCQ (BASED ON PRACTICAL)	
5.	C) Green color of salt fades and a gas with the smell of burning sulfur is evolved.	1M
6.	b) Metals are malleable and ductile.	1M
7.	b) HCl	1M
8.	d) B and D	1M
9.	BIO- MCQ	
10.	BIO- MCQ	
11.	BIO- VSA (BASED ON PRACTICAL)	
12.	BIO- VSA (BASED ON PRACTICAL)	
13.	BIO- VSA	
14.	BIO- ASSERTION/REASONING TYPE	
15.	PHY- VSA (BASED ON PRACTICAL)	
16.	PHY- VSA	
17.	PHY- ASSERTION/REASONING TYPE	
18.	A white precipitate is formed due to the formation of $BaSO_4$	½ + ½ = 1M
19.	Substance oxidized – CO , Substance Reduced - Fe_2O_3	½ + ½ =1M
20.	Due to evolution of CO_2 gas.	1M
21.	PHY OR	
22.	PHY	
23.	PHY OR	
24.	(A) Alkai, Eg. NaOH or Name (B) Bee sting contains Methanoic acid. Baking Soda being basic in nature, neutralizes acid and gives relief. (C) Strong – H_2SO_4 , HNO_3 (Any one from each) Weak – CH_3COOH , H_2CO_3 OR (A) Soil is acidic (B) i) A is $CuSO_4.5H_2O$ or (Hydrated copper suphate) B is water of crystallization. C is $CuSO_4$ or Anhydrous copper sulphate D is water ii) Any example of a hydrated salt (C) Plaster of Paris absorbs moisture and changes to gypsum, a hard mass.	½ + ½ = 1M ½ + ½ = 1M ½ + ½ = 1M ½ M 1/4x4= 1M ½ M 1M

25.	i) Reddish – copper , greyish – Silver ii) Corrosion iii) Green – Copper carbonate or $CuCO_3$ Black - Silver sulphide Or AgS	½ + ½ = 1M 1 M ½ + ½ = 1M
26.	(A) Al is reducing agent. Al is more reactive, because it displaces Mn from MnO_2. (b) Because it undergoes photochemical reaction in the presence of sunlight. OR (A) i) Endothermic ii) Exothermic (B) i) $Cu + 2AgNO_3 \rightarrow Cu(NO_3)_2 + 2Ag$ Displacement Reaction ii) $3H_2 + N_2 \rightarrow 2NH_3$ Combination reaction	1M ½ + ½ = 1M 1M ½ + ½ = 1M ½ x 4 = 2M
27.	BIO OR	
28.	BIO	
29.	BIO	
30.	BIO	
31.	PHY OR	
32.	PHY	
33.	(A) i) By heating alone the metal oxide ii) Electrolytic Reduction/ Electrolysis of its molten ore. iii) By using suitable reducing agents like carbon (coke) or CO or by using highly reactive metals like sodium , aluminium as reducing agents. (B) Formation of Na_2O by transfer of electrons. (C) One difference OR (A) Any two differences with chemical equations. (B) Electrolytic Refining. Pure copper is used as cathode. Impure copper is used as anode. Copper sulfate as electrolyte. (C) Correct definition with one example.	3x1= 3M 1M 1M 1 + 1 = 2M ½ x 4 = 2M 1M
34.	(A) Three balanced chemical equations. (B) Solution A has max. concentration of H+ ions. A is acidic, while B and C are Basic in nature	3 x 1= 3M ½ x 4 = 2M
35.	BIO OR	
36.	BIO	